
GASES

The kinetic theory of gases

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The kinetic theory and gases

The kinetic theory considers that gases are made up of very small particles with no binding forces. These particles move freely inside the container, in a straight line. They only change direction when they collide with each other or with the walls of the container.

Pressure in a gas is a measure of the number of collisions of the particles against the walls of the container. More collisions means more pressure.

Temperature in a gas is proportional to the speed of the particles.

Absolute zero

Temperature is a measure of the speed of the particles in the gas. The lower the temperature, the less speed particles will have.

When particles don't move, the gas is at absolute zero temperature. Temperature cannot decrease any further.

This lowest temperature is called absolute zero and is equivalent to $-273.15\text{ }^{\circ}\text{C}$.